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help you figure out? 2. What are all
reactions dependent upon? 3. If I have
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oxygen molecules, how many molecules
of water can I make?

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Answer Key Chapter 12: Stoichiometry Mole Ratios Questions 1. Aluminum reacts with oxygen to produce aluminum oxide as follows: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
 a. If you use 2.3 moles of Al, how many moles of Al_2O_3 can you make? b. If you want 3.9 moles of Al_2O_3 , how many moles of O_2 are needed? 2. In the presence of sulfuric acid, metallic iron forms iron(III) sulfate: $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$...

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Stoichiometry 379 CHAPTER 12 Assessment 36 Chapter 12 stoichiometry

test b answer key. a. Two formula units KClO_3 decompose to form two formula units KCl and three molecules O_2 . b. Chapter 12 stoichiometry test b answer key. Four molecules NH_3 react with six molecules NO to form five mol-

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