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$aA + bB \xrightarrow{C} X$ with a catalyst C. Rate = $k[A]^m[B]^n[C]^p$. The exponents m, n, and p •are the reaction order •can be 0, 1, 2 or fractions •must be determined by experiment! 12. © 2009 Brooks/Cole - Cengage. Interpreting Rate Laws. Rate = $k[A]^m[B]^n[C]^p$. •If m = 1, rxn. is 1st order in A Rate = $k[A]^1$.

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When tin comes in contact with the oxygen in the air, tin ...

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