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Determine the products of the following reaction, then ...

Since the "possible" products of the double displacement reaction do not precipitate in an aqueous solution, these compounds will simply dissociate in the solution. If we write the complete ionic...

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a reactant in aqueous solution is indicated with (aq); formation of a solid is indicated with (s) Describe what happens when an ionic substance is dissolved in water the ions separate, become hydrated, and behave independently What is meant by a strong electrolyte?

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substances as acids and bases that would not be classified as such according to the Arrhenius definitions.

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carried out in solution and do not take place at all if the solid reactants are simply. mixed. The reaction of mercury (II) acetate with sodium iodide. When colorless aqueous solutions of each reactant are. mixed, they produce a red precipitate, mercury (II)

iodide, which is the result of an exchange reaction.

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In your textbook, read about aqueous solutions, reactions that form precipitates, reactions that form water, and reactions that form gases. Circle the letter of the choice that best completes the statement or answers the question. 1. A spoonful of sodium chloride is dissolved in a liter of water.