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CHAPTER 14. SOLIDS, LIQUIDS AND GASES. STATES OF MATTER Solids - have a definite shape and volume. WHY? Particles of a solid are packed close

together Particles only vibrate in place  
CRYSTAL SOLIDS Particles are arranged in a repeating pattern EX. Diamonds, emeralds, ruby, salt, quartz Liquids Definite volume, no set shape Particles have enough space to move around one another allows a liquid to flow from place to place FLUID - "substance that flows" Liquids also have - VISCOSITY

## CHAPTER 14 SOLIDS, LIQUIDS AND GASES

Chapter 14: Liquids and Solids; Chapter 16: Acid/Base Chemistry; Important. 3 States of Matter: 1. Solid: rigid; has a fixed shape and volume ex. Ice-water molecules are locked into rigid positions and are close together 2. Liquid: has a definite volume but takes the shape of its container

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1. Solid: Ice molecules are locked into rigid positions and are close together 2. Liquid: Water molecules are still close together but can move around to some extent 3. Gas: Steam molecules are far apart and move randomly

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Chapter 14: Liquids and Solids; Chapter 15; Chapter 16: Acid/Base; Liquids and Solids. 3 States of Matter: 1. Solid: rigid; has a fixed shape and volume ex. Ice-water molecules are locked into rigid positions and are close together 2. Liquid: has a definite volume but takes the shape of its container

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Chapter 14: Liquids and Solids. a covalent bond in which the shared electrons are shared equally between

the two atoms involved in the bond. diatomic molecules. a covalent bond in which the shared electrons are shared unequally between the two atoms involved in the bond; the electrons spend more time revolving around the more electronegative atom.

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Dissolution of Solids in Liquids • The energy released (exothermic) when a mole of formula units of a solid is formed from its constituent ions (molecules or atoms for nonionic solids) in the gas phase is called the . crystal lattice energy. • The crystal lattice energy is a measure of the attractive forces in a solid.

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Chapter 14: Liquids and Solids. Lesson Plans Chapter 14: Phases Changes. Chapter 14: Liquids and Solids. I. Phases of matter and phase changes a. Recall the three main phases of matter: Energy is involved during the transition from one phase of matter to another. You should learn how energy is involved in each of these phase transitions, and you should be able to name each of these phase transitions.

### **Chapter 14: Liquids and Solids**

49 Terms. a02112003. chapter 14:liquids & solids. condensed state of matter consists of..... intermolecular forces (IMF'S) 3 types of IMF'S (weakest to strongest) dispersion forces (London dispersion fo.... solids and liquids. the outer forces that make molecules stick together to form li....

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Section 14.2 Energy Requirements for the Changes of State Phase Changes • Solid to Liquid As energy is added, the motions of the molecules increase, and they eventually achieve the greater movement and disorder characteristic of a liquid. • Liquid to Gas As more energy is added, the gaseous state is eventually reached, with the individual

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CHAPTER 14: LIQUIDS AND SOLIDS INTRODUCTION This chapter discusses the properties of liquids and solids. You will learn what makes the particles in solids stay together and why some liquids boil at higher temperatures than others. One liquid of particular importance is water. Liquid water is one of the most important parts of

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The solid is less dense than the liquid. Liquids- Particles flow. Fluidity- has the

ability to flow. Both Gases and Liquids are classified as fluids. Take the shape of container. Has a definite volume. Denser than gases. Viscosity - Measure of the resistance of a liquid to flow. Water is not very viscous but molasses is. "slow as molasses"

### **Chapter 13 Notes Liquids and Solids**

14.1: Prelude to Solids and Liquids.

Liquids flow when a small force is placed on them, even if only very slowly. Solids, however, may deform under a small force, but they return to their original

shape when the force is relaxed.

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