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# The Way Is Shut

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*The Way Is Shut*

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by Marco & Bond

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### Section 3 Predicting The Products

Section Three: Predict the products for the following reactions. Draw their mechanisms using arrow-pushing notation. Show the major and the minor product, and show any stereochemistry and intermediates as appropriate. More than one type of reaction (but not more than two) may be possible for each. Each problem is worth 25 points.

### Section Three: Predict The Products For The Follow ...

To predict the products: Replace the single element with the corresponding element in the compound !A metal will replace the cation in the compound !A non-metal will replace the anion in the compound If the element is more reactive than the one it attempts to replace, then the reaction will be spontaneous and exothermic

### Steps to Predicting the Products of Chemical Reactions

Question: Section Three: Predict The Products For The Following Reactions. Draw Their Mechanisms Using Arrow-pushing Notation. Show The Major And The Minor Product, And Show Any Stereochemistry And Intermediates As Appropriate. More Than One Type Of Reaction (but Not More Than Two) May Be Possible For Each. Each Problem Is Worth 25 Points.

### Solved: Section Three: Predict The Products For The Follow ...

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Section 3: Predicting the products of chemical reactions

### Section 3: Predicting the products of chemical reactions

Predicting+WS - Section 3 Predicting the products of ...  $2 \text{KClO}_3 (\text{s}) \rightarrow 3 \text{O}_2 (\text{g}) + 2 \text{KCl} (\text{s})$  The potential products in double-replacement reactions are simple to predict; the anions and cations simply

### Section 3 Predicting The Products Of Chemical Reactions

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### Section 3 Predicting The Products Of Chemical Reactions

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Unit 1.3: Predicting Products of Chemical Reactions. Look at the reactants of a chemical reaction and predict the products. f you got the correct compounds but in a different order than was the correct answer then use the "I was RIGHT" button. 3 things to always remember when balancing chemical equations: 1)  $\text{HX} + \text{Y} (\text{OH}) \rightarrow \text{XY} + \text{H} (\text{OH})$  No matter what elements X and Y are; any double replacement reaction with a hydroxide bonding to a hydrogen can be written as  $\text{H}_2\text{O} * \text{OR} * \text{H} (\text{OH})$ .

### Unit 1.3: Predicting Products of Chemical Reactions ...

Answers for Predicting Products of Chemical Reactions For all combustion reactions of hydrocarbons or hydrogen-carbon-oxygen molecules, the products will always be  $\text{CO}_2$  and  $\text{H}_2\text{O}$  (assuming a complete reaction). a. Combustion:  $\text{C}_6\text{H}_{12} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$  b. Combustion:  $2\text{C}_4\text{H}_6 + 11\text{O}_2 \rightarrow 8\text{CO}_2 + 6\text{H}_2\text{O}$  c. Combustion:  $\text{C}_6\text{H}_{10}\text{O}_3 + 7\text{O}_2 \rightarrow 6\text{CO}_2 + 5\text{H}_2\text{O}$  ...

### Answers for Predicting Products of Chemical Reactions

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Also, this is an acid-base reaction, so the products should be salt and water. 11. Decomposition  $\text{Al}_2(\text{CO}_3)_3 + \text{heat} \rightarrow \text{Al}_2\text{O}_3 + 3\text{CO}_2$ . Another decomposition that will generate  $\text{CO}_2$  and a metal oxide (note that Al forms a +3 ion,  $\text{Al}^{3+}$ , and monatomic oxygen will have a -2 charge,  $\text{O}^{2-}$ ). 12. Synthesis  $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$

### Answers for Predicting Products of Chemical Reactions

Section 3: Predicting the products of chemical reactions Predict the products of the following reactions: 2) 5) 6) 8) 9) 10)  $\text{Ag} + \text{CuSO}_4 \rightarrow$  NC.. Type: Type:  $\text{O}_2 +$  Type: Type:  $\text{BaSO}_4 + \text{H}_2\text{O} +$  LiBr + Type:  $\text{CuSO}_4$  Type: A Type: Type: Type: All Rights Reserved WKS001x002 © 2000 Cavalcade Publishing (<http://mvw.cavalcadepublishing.com>)

### Nicolet High School

Look at the reactants of a chemical reaction and predict the products. For this set you will have to balance the equations after you determine the products. Terms in this set (30)  $\text{C}_4\text{H}_6 + \text{O}_2 \rightarrow$

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Section 3 Predicting The Products Of Chemical Reactions Look at the reactants of a chemical reaction and predict the products. f

you got the correct compounds but in a different

### Section 3 Predicting The Products Of Chemical Reactions

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Why we need to predict reaction products As with many chemistry-oriented things, the skill of predicting reaction products is useless if you don't actually use it for something. Actually, that's true for just about everything, whether chemistry-oriented or not.

### Predicting reaction products | The Cavalcade o' Chemistry

Section 2: Predicting the products of chemical reactions • Predict the products of the following reactions. • Predict the states of matter for double replacement reactions. • Finally balance each reaction. 1)  $\text{Ag} + \text{CuSO}_4 \rightarrow$  2)  $\text{NaI} + \text{CaCl}_2 \rightarrow$  3)  $\text{O}_2 + \text{H}_2 \rightarrow$  4)  $\text{HCN} + \text{CuSO}_4 \rightarrow$  5)  $\text{HNO}_3 + \text{Fe}(\text{OH})_3 \rightarrow$  6)  $\text{LiBr} + \text{Co}(\text{SO}_3)_2(\text{aq}) \rightarrow$  7)  $\text{LiNO}_3 + \text{Ag} \rightarrow$  8)  $\text{H}_2\text{CO}_3 \rightarrow$  9)  $\text{H}_2\text{SO}_4 + \text{NH}_4\text{OH} \rightarrow$  10)  $\text{C}_4\text{H}_8 \rightarrow$

### PLEASE HELP! Chemistry HW predicting products of chemical ...

Reaction Products Worksheets: (Updated 3-30-16) Double Displacement Reactions: What will a double displacement reaction form? Will it take place at all? With this worksheet and a solubility table, they'll understand in no time! Reaction Products Worksheet: Your students will practice making predictions about the products of chemical reactions, given only the reagents.

### Predicting Reaction Products | The Cavalcade o' Teaching

3 Types of Chemical Reactions Notes • Synthesis - two or more

elements or compounds combine to form one compound. •  
 Decomposition-a single compound decomposes into two or more  
 elements or smaller compounds. • Single Replacement - a metal  
 will replace a less active metal in an ionic compound OR a  
 nonmetal will replace a less active nonmetal. • Double  
 Replacement - the metals in ionic ...

### Unit 9 Chemical Equations and Reactions

Putting these predictions together, the product is likely to be  $\text{AlCl}_3$ .  
 In fact, if aluminum metal and chlorine gas are allowed to  
 react, solid  $\text{AlCl}_3$  is the predominant product.  $2 \text{Al (s)} + 3 \text{Cl}_2 \text{ (g)}$   
 $\rightarrow 2 \text{AlCl}_3 \text{ (s)}$  The synthesis reaction involving the non-metals  
 hydrogen gas and bromine can be approached similarly.