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Chapter 15 - Acids and Bases - Questions and Problems ...

Chapter 15: Acids and BasesAcids and Bases. Arrhenius Definitions: acids- compounds that produce an increase in [H+] when dissolved in water. bases- compounds that produce an increase in [OH-] when dissolved in water Lewis Definitions: acids- electron pair acceptors.

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An acid-base reaction occurs in aqueous solution between hydro- chloric acid, a strong acid that completely dissociates to produce H. 3O+, and sodium hydroxide, a strong base that completely dissociates to produce OH-. The formula equation for this reaction is written as follows. HCl(aq) +NaOH(aq) → NaCl(aq) +H.

CHAPTER 15 Acids and Bases

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Acid and Base Worksheet - Answers. 1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following acid-base reactions and indicate each conjugate acid-base pair: a) HNO3 + OH-1 (H2O + NO3-1. HNO3 and NO3-1 make one pair OH-1 and H2O make the other. b) CH3NH2 + H2O (CH3NH3+ + OH-1

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Chapter 15 Acids and Bases Multiple Choice Questions 1) ____ is found in carbonated beverages due to the reaction of carbon dioxide with water. A) CH3COOH; B) H2CO3; C) HCOOH; D) C6H5COOH; E) CH3CH2COOH; Answer: B. Diff: 1 Page Ref: 15.2 2) Which of the following species is amphoteric? A) CO32-B) HF; C) NH4+ D) HPO42-

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The reaction is as follows: Z n O (a q) + 2 H C l (a q) → Z n C l 2 (a q) + H 2 O (l) Zinc oxide reacts with sodium hydroxide to produce zincate sodium and water. This reaction takes place at a temperature of 500-600°C.it is exothermic reaction. ZnO + 2 NaOH → Na 2 ZnO 2 + H 2 O. g. Limestone is calcium carbonate.