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# Rapport De Stage Architecte

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Chapter 9 Chemical Names and Formulas83 SECTION 9.3 NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS (pages 268-270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268-269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

### Name Date Class CHEMICAL NAMES AND FORMULAS 9

SECTION 9.2 NAMING AND WRITING FORMULAS FOR IONIC COMPOUNDS I. Write the formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas for the compounds formed c. b. f. 3. Name the following inary ionic compounds. sodium sulfide g. CuCl: h. snc.l. a. MnO2 c. d. SrBr2 221 Chapter 9 Chemical Names and Formulas

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Chapter 9 Chemical Names and Formulas 9.1 Naming Ions 9.2 Naming and Writing Formulas for Ionic Compounds 9.3 Naming and Writing Formulas for Molecular Compounds 9.4 Naming and Writing Formulas for Acids and Bases 9.5 The Laws Governing How Compounds Form

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Chapter 9 "Chemical Names and Formulas" ... in parenthesis after the name of the metal (Table 9.2, p.255) Predicting Ionic Charges Some of the post-transition elements also have more than one possible oxidation state. Tin (II) =  $\text{Sn}^{2+}$  Lead (II) =  $\text{Pb}^{2+}$  Tin (IV) =  $\text{Sn}^{4+}$  Lead (IV) =  $\text{Pb}^{4+}$

### "Chemical Names and Formulas"

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Chemical Names and Formulas 281 CHAPTER 9 Assessment 42. a.  $2^{-}$  b.  $1^{+}$  c.  $1^{-}$  d.  $3^{+}$  43. a.  $2^{+}$  b.  $2^{+}$  c.  $3^{+}$  d.  $1^{+}$  44. a. barium ion b. iodide ion c. silver ion d. mercury(II) ion 45. cyanide,  $\text{CN}^{-}$  and hydroxide,  $\text{OH}^{-}$  46. a. hydroxide ion b. lead(IV) ion c. sulfate ion d. oxide ion 47. zero; A compound is

electrically neutral. 48. The symbols for the cation and

### CHAPTER 9 Study Guide - Quia

Chapter 9 Practice Problems: Chemical Names and Formulas. Section 9.1: Naming Ions. 1. What is the charge on the ion typically formed by each element? a. oxygen c. sodium e. nickel, two electrons lost . b. iodine d. aluminum f. magnesium. 2. How many electrons does the neutral atom gain or lose when each ion forms? a.  $\text{Cr}^{3+}$  c.  $\text{Li}^{+}$  e.  $\text{Cl}^{-}$  b.  $\text{P}^{3-}$  ...

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