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ARITHMETIC OF EQUATIONS (pages 353–358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process.

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Guided Practice: Stoichiometry Mass to Mass Problems To convert from mass in grams of a reactant to volume, in liters, of a product (reverse the process for liters to grams):

- Use factor label method
- Use mass of reactant from the Periodic Table 1 mol = ____ g
- Use the mole to mole ratio from the balanced reaction

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