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summarized in the  
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law :  $PV = nRT$ . where  
 $n$  is the number of  
moles of gas, and  $R =$   
 $8.31 \text{ J / (mol K)}$  is  
known as the universal  
gas constant.

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**AP Physics B - Gas**  
**laws**

There are various  
versions of the gas  
laws and some are the

same version of the others but said in a slightly different way:

1. All molecules of a particular gas are identical. 2. The internal energy of the gas is entirely kinetic. 3. All collisions between molecules and the walls of the container are completely elastic. 4. Newton's laws of motion apply.

### **Gas Laws - Physics A-Level - Physics A-Level**

When the only properties of an ideal gas that are changing are volume and temperature, we use Charles' Law (a derivative of the Ideal Gas Law). Charles' Law is as follows: We're given all but the new volume,

### **Ideal Gas Law - AP**

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### **Ap Physics B Gas**

#### **Laws -**

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Boyle's law—named for Robert Boyle—states that, at constant temperature, the pressure  $P$  of a gas varies inversely with its volume  $V$ , or  $PV = k$ , where  $k$  is a constant. Charles's law—named for J.-A.-C. Charles (1746–1823)—states that, at constant pressure, the volume  $V$  of a gas is directly proportional to its absolute (Kelvin) temperature  $T$ , or  $V/T = k$ .

### **gas laws | Definition & Facts | Britannica**

The Ideal Gas Equation

The gas laws can be combined into a general equation that describes the physical behavior of all gases.

11.5 Boyle's law

Avogadro's law

Charles's law  $PV = nRT$

rearrangement  $R$  is the proportionality

constant, called the

gas constant. B. Ideal

Gas Law UNIVERSAL

GAS CONSTANT  $R =$

$8.3145 \text{ J/mol}\cdot\text{K}$

$R=0.0821 \text{ L atm/mol K}$

$PV=nRT$   $R = 0.0821$

$\text{liter}\cdot\text{atm/mol}\cdot\text{K}$   $R =$

$8.3145 \text{ J/mol}\cdot\text{K}$   $R =$

$8.2057 \text{ m}^3\cdot\text{atm/mol}\cdot\text{K}$

$R = 62.3637$

$\text{L}\cdot\text{Torr/mol}\cdot\text{K}$  or

$\text{L}\cdot\text{mmHg/mol}\cdot\text{K}$  Gases

only occupy about 0 ...

### **PPT: GAS LAWS - The colony HIGH SCHOOL**

The gas in a cylinder

has a pressure of 210kPa at a temperature of 27°C. Calculate the new pressure when the temperature of the gas rises to 81°C.  $27^{\circ}\text{C} = (27+273)\text{K} = 300\text{K}$ .  $81^{\circ}\text{C} = (81+273)\text{K} = 354\text{K}$ .  $P_1 / T_1 = P_2 / T_2$   $210 / 300 = P_2 / 354$ .  $P = 210 \times 354 / 300 = 247.8 \text{ kPa}$

### **PHYS - Questions on Gas Laws - physics gas laws**

The pressure of a gas is directly proportional to its temperature when volume is constant.

Symbolically...  $P \propto T$  (V constant) An isochoric process is one that takes place without any change in volume. This relationship doesn't really have a name, but I have heard it called the "pressure law" or (mistakenly)

"Gay-Lussac's law".

### **Gas Laws - The Physics Hypertextbook**

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Gas laws and the kinetic model The behaviour of gases is described in terms of the kinetic theory, which considers the motion of molecules in the gas. The gas law links pressure, volume

and...

### **Volume and temperature of a gas - Gas laws and the kinetic ...**

The behavior of an ideal gas, that is, the relationship of pressure (P), volume (V), and temperature (T), can be summarized in the ideal gas law: Ideal gas law :  $PV = nRT$ . where n is the number of moles of gas, and  $R = 8.31 \text{ J / (mol K)}$  is known as the universal gas constant. An alternate way to express the ideal gas law is in terms of N, the number of molecules, rather than n, the number of moles. N is simply n multiplied by Avogadro's number, so the ideal gas law can be written as:

### **Ideal Gas Law |**

### **CourseNotes**

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ideal gas law 1 point .  
 $pVnRT = (200 \text{ N/m}^2 \cdot 20 \text{ m}^3 \cdot 1 \text{ mol} \cdot 8.32 \text{ J/mol K} \cdot 23) / ( ) = ( ) ( ) T$

### **Thermodynamics**

#### **Free Response**

#### **Questions KEYIdeal Gas Law**

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