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$4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$ ΔH_f° of $\text{NH}_3 = -46.0 \text{ kJ mol}^{-1}$ ΔH_f° of $\text{NO} = +90.5 \text{ kJ mol}^{-1}$ ΔH_f° of $\text{H}_2\text{O} = -242 \text{ kJ mol}^{-1}$
 $\Delta H^\circ_{\text{reaction}} = 4y + [6 \times (-242)] - [4 \times (-46.0)]$ (1) [1] $\Delta H^\circ_{\text{reaction}} = 4y - 1452 + 184$ $\Delta H^\circ_{\text{reaction}} = -906 \text{ kJ mol}^{-1}$
 $-1 \times 4y = -906 + 1452 - 184 = 362$ (1) [1] $y = +90.5 \text{ kJ mol}^{-1}$

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