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Ideal gas law: $\rho = \frac{p}{RT} = \frac{124.8 \text{ kPa}}{1 \text{ kg/K} \cdot 2077 \text{ J/K}} = 1.293 \text{ kg/m}^3$
 $1000 \text{ Pa} = 1 \text{ kPa}$
 $\text{J/Nm} = \text{Pa} \cdot \text{m}^3 = 0.181 \text{ kg/m}^3$
b.) Situation: A sphere of 93 mm diameter contains an ideal gas. $T = 20^\circ\text{C} = 293.2 \text{ K}$
Find: Calculate the density of argon at a vacuum pressure of 8.8 psi.
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