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Chapter 11: Chemical Reactions

I II III I. Intro to Reactions Ch. 11 - Chemical Reactions A.Signs of a Chemical Reaction Evolution of heat and light Formation of a gas Formation of a precipitate Color change B.Law of Conservation of Mass mass is neither created nor destroyed in a chemical reaction $4\text{ H}_2 + \text{O}_2 \rightarrow 2\text{ H}_2\text{O}$ 4 g 32 g 36 g total mass stays the same atoms can only rearrange

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As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano. Heat, light, and gas are produced as a large pile of fluffy green chromium (III) oxide forms. We can describe this reaction with a chemical equationAn expression that gives the identities and quantities of the substances in a chemical reaction.

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A combustion reaction is any reaction that burns in the presence in oxygen gas (O₂). The products are usually an oxide of the element. If there is the combustion of a hydrocarbon, carbon dioxide and water vapor are products. CH 11- Stoichiometry

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Chapter 11: Properties of Reactions. An oxidation number is a positive or negative number that is assigned to an atom to indicate its degree of oxidation or reduction. The term oxidation state is often used interchangeably with oxidation number. A partial electron transfer is a shift in the electron density near an atom as a result of a change in the other atoms to which it is covalently bonded.

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330 Chapter 11 11.2 Types of Chemical Reactions Often charcoal briquettes provide the heat for barbeque grills through the burning of carbon. Have you ever felt the heat and smelled the smoke coming from a burning charcoal grill? The heat and smoke are the products of a combustion reaction. Combustion is one of the five general types of chemical